# Standard Reduction Potentials

| half-reaction | $n$ | $E^o$ (V) |
|---|---|---|
| $F_2(g) + 2e^- \rightarrow 2F^-(aq)$ | 2 | 2.866 |
| $H_2O_2(aq) + 2H^+(aq) \rightarrow 2H_2O(l)$ | 2 | 1.776 |
| $PbO_2(s) + SO_4{}^{2-}(aq) + 4H^+(aq) + 2e^- \rightarrow PbSO_4(s) + 2H_2O(l)$ | 2 | 1.6913 |
| $MnO_4{}^-(aq) + 4H^+(aq) + 3e^- \rightarrow MnO_2(s) + 2H_2O(l)$ | 3 | 1.673 |
| $MnO_4{}^-(aq) + 8H^+(aq) + 5e^- \rightarrow Mn^{2+}(s) + 4H_2O(l)$ | 5 | 1.507 |
| $Cl_2(g) + 2e^- \rightarrow 2Cl^-(aq)$ | 2 | 1.3583 |
| $O_2(g) + 4H^+(aq) + 4e^- \rightarrow 2H_2O(l)$ | 4 | 1.229 |
| $Ag^+(aq) + e^- \rightarrow Ag(s)$ | 1 | 0.7996 |
| $Fe^{3+}(aq) + e^- \rightarrow Fe^{2+}(aq)$ | 1 | 0.770 |
| $O_2(g) + 2H^+(aq) + 2e^- \rightarrow H_2O_2(aq)$ | 2 | 0.68 |
| $Cu^{2+}(aq) + 2e^- \rightarrow Cu(s)$ | 2 | 0.342 |
| $AgCl(s) + e^- \rightarrow Ag(s) + Cl^-(aq)$ | 1 | 0.2223 |
| $2H^+(aq) + 2e^- \rightarrow H_2(g)$ | 2 | 0.000 |
| $Fe^{2+}(aq) + 2e^- \rightarrow Fe(s)$ | 2 | –0.447 |
| $Li^+(aq) + e^- \rightarrow Li(s)$ | 1 | –3.05 |